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contents:

thermodynamics .

chapter 01:

thermodynamic

properties and state of  
pure substances.

chapter 02: work and

heat. chapter 03:

energy and the first  
law of

thermodynamics.

chapter 04: entropy

and the second law of  
thermodynamics.

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chapter 05:  
irreversibility and  
availability

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Practice Problems  
Thermodynamics. 1.  
Why is the entropy  
change in a system not  
always a reliable  
predictor of whether  
the process producing  
the change is  
spontaneous? ...

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Activities are approximated by using solution concentrations in units of molarity (divided by 1 M to remove the units) and gas partial pressures in units of atm (divided by 1 ...

## **CHM 112 Thermodynamics Practice Problems Answers**

Problem : Given that the free energy of formation of liquid

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water is  $-237 \text{ kJ / mol}$ , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate  $\Delta G$  for the reaction, which is  $-2 ( -237 \text{ kJ / mol} ) = 474 \text{ kJ / mol}$ . Knowing that  $\Delta G = -nFE^{\circ}$  and  $n = 4$ , we calculate the potential is  $-1.23 \text{ V}$ .

## **Thermodynamics: Problems and Solutions |**



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From first law of  
Thermodynamics  
 $\Delta U = \Delta Q - \Delta W$  Since  
 $\Delta U = 0$   $\Delta Q = \Delta W$  Also

$PV = nRT$  As  $T$  is  
constant  $PV = \text{constant}$

Question-.2 Two  
absolute scales A and  
B have triple points of  
water defined as  $200A$   
and  $350A$ . what is the  
relation between  $T_A$   
and  $T_B$  Solution-2

Given that on absolute  
scale Triple point of  
water on scale A = 200

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1 - Introduction: Basic  
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Thermodynamics ... In  
many courses, the  
instructor posts copies  
of pages from the  
solution manual. Often  
the solution manual  
does little more than  
show the quickest way

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to obtain the answer and says nothing about WHY each step is taken or HOW the author knew which step to ...

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### **Thermodynamics - Example Problems**

Thermodynamics - problems and solutions. The first law of thermodynamics. 1. Based on graph P-V below, what is the ratio of the work done by the gas in the process

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I, to the work done by the gas in the process II? Known : Process 1 : Pressure (P) = 20 N/m<sup>2</sup>  
2. Initial volume (V<sub>1</sub>) = 10 liter = 10 dm<sup>3</sup> = 10 x 10<sup>-3</sup> m<sup>3</sup>

## **Thermodynamics - problems and solutions | Solved Problems ...**

The following are common thermodynamic equations and sample problems showing a

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situation in which each might be used.

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*Page 13/29*

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The first law of thermodynamics - problems and solutions. 1. 3000 J of heat is added to a system and 2500 J of work is done by the system. What is the change in internal energy of the system?  
Known : Heat (Q) = +3000 Joule. Work (W) = +2500 Joule .  
Wanted: the change in internal energy of the

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system Solution :  
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**The first law of  
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Practice:

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questions. This is the  
currently selected  
item. Thermodynamics  
article. ... First law of  
thermodynamics. First  
law of thermodynamics  
problem solving. PV  
diagrams - part 1:  
Work and isobaric

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processes. PV diagrams - part 2: Isothermal, isometric, adiabatic processes. Second law of thermodynamics. Next lesson ...

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Mechanical - Engineering  
Thermodynamics - The Second Law of Thermodynamics 1.  
Two kg of air at



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500kPa, 80°C expands adiabatically in a closed system until its volume is doubled and its temperature becomes equal to that of the surroundings which is at 100kPa and 5°C.

## **Solved Problems: Thermodynamics Second Law**

SOLUTIONS

THERMODYNAMICS

PRACTICE PROBLEMS

FOR NON-TECHNICAL

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Properties 1. If an object has a weight of 10 lbf on the moon, what would the same object weigh on Jupiter? Jupiter...

## **Thermodynamic Properties**

Physics problems:  
thermodynamics. Part  
1 Problem 1. A rapidly  
spinning paddle wheel  
raises the temperature  
of 200mL of water from

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21 degrees Celsius to 25 degrees. How much a) work is done and b) heat is transferred in this process? Solution .  
Problem 2. The temperature of a body is increased from  $-173\text{ C}$  to  $357\text{ C}$ .

## **Physics Problems: Thermodynamics**

Practice Problems  
Chapter 5: 1, 3, 4 (3  
problems Practice  
Problem Solutions)  
Homework Problems

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Chapter 5: 8, 13, 15 (3  
problems ans) Test

Yourself P. 248

Problem 1; Practice

Problems Chapter 6: 1;

(2 problems Practice

Problem Solution) Test

Yourself P. 236;

Problems Chapter 6:

4,6,7,8,11 (5 problems,

ans) Part of Chap. 4,

End of Chap. 4

**Chemical  
Engineering  
Thermodynamics**

Solved Problems on Th

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ermodynamics:-Problem 1:-A container holds a mixture of three nonreacting gases:  $n_1$  moles of the first gas with molar specific heat at constant volume  $C_{v1}$ , and so on. Find the molar specific heat at constant volume of the mixture, in terms of the molar specific heats and quantities of the three separate gases.

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634 Heat Engines,  
Entropy, and the  
Second Law of  
Thermodynamics  
SOLUTIONS TO  
PROBLEMS Section  
22.1 Heat Engines and  
the Second Law of  
Thermodynamics P22.1  
(a)  $e = \frac{W}{Q_h} = \frac{Q_c}{Q_h} = \frac{Q_h - W}{Q_h} = 1 - \frac{W}{Q_h} = 1 - \frac{Q_c}{Q_h} = \frac{Q_h - Q_c}{Q_h} = \frac{W}{Q_h} = \eta_{eng}$  J

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360 J 25 0 00694.. or  
694%. (b)  $Q_c = -$

$= - = \text{eng } 360 \text{ } 25 \text{ } 0$   
335 J J J . P22.2

$W = Q_c = - = hc200 \text{ J}$   
(1)  $e = W = Q_c = h = Q_c = h$

$= - = \text{eng } 10300$ . (2)

From (2),

$Q_c = 0700$ . (3) Solving

(3) and (1)

simultaneously,

## **Heat Engines, Entropy, and the Second Law of Thermodynamics**

- So far you've seen



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the First Law of Thermodynamics. This is what it says. Let's see how you use it. Let's look at a particular example. This one says, let's say you've got this problem, and it said 60 joules of work is done on a gas, and the gas loses 150 joules of heat to its surroundings.

**First law of  
thermodynamics  
problem solving**

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