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Answer the following questions with regards to the Heating Curve. 1) What is happening to the average potential energy of the molecules in the sample during section 3? Potential Energy remains constant. It only increases or decreases when temperature is remains constant on the plateaus. ...

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AP CHEMISTRY REVIEW WORKSHEET

(Unit 4 - Thermochemistry) 1. $3 \text{SO}_2 (\text{g}) + 2 \text{HNO}_3 (\text{aq}) + 2 \text{H}_2\text{O} (\text{l}) \rightarrow 3 \text{H}_2\text{SO}_4 (\text{aq}) + 2 \text{NO} (\text{g})$ a. Calculate ΔH for the above reaction. Based on the enthalpy change, determine whether the reaction is endothermic or exothermic and whether it would be thermodynamically

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Problem #1: A cubic block of uranium metal (specific heat = $0.117 \text{ J g}^{-1} \text{ } ^\circ\text{C}^{-1}$) at $200.0 \text{ } ^\circ\text{C}$ is dropped into 1.00 L of deuterium oxide (also called "heavy water") (specific heat = $4.211 \text{ J g}^{-1} \text{ } ^\circ\text{C}^{-1}$) at $25.5 \text{ } ^\circ\text{C}$.

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AP Chemistry Practice Test, Ch. 6:

Thermochemistry Name _____ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure.

A) endothermic, positive

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AP Chemistry Practice Test, Ch. 6: Thermochemistry ...

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True or False If the answer is false, you must explain why. 1. _____ Chemical energy is a form of kinetic energy. 2.

_____ When we mix water at 100 C with water at 20 C the final water temp. is

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always 60 C. 3. _____ When we put a 100g piece of metal at 100 C into 100g of water at 20 C the final temp is 20 C. 4.

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PLAY. Match. Gravity. Created by.
Dragonknighted1. Terms in this set (79)
The _____ is the set of properties that
describe the system. state. The _____ is
comprised of all the matter in the
reaction. system

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[FREE] Chapter 6 Thermochemistry Review Answers. Thermochemistry Test Preview Matching Match each item with the correct statement below. a. calorimeter d. enthalpy b. calorie e. specific heat c. joule f. heat capacity ____

1. quantity of heat needed to raise the temperature of 1 g of water by 1 C
2. ____ SI unit of energy
3. ____ quantity of

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heat needed to change the temperature of 1 g of a ...

Chapter 6 Thermochemistry Review Answers

Thermochemistry Exam1 and Problem Solutions 1. Which ones of the following reactions are endothermic in other words ΔH is positive? I. $\text{H}_2\text{O}(\text{l}) +$

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10,5kcal → H₂O(g) ΔH₁ II. 2NH₃ +22kcal

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_____ 1) How many joules are required to raise 20.0 g of ice from -10.0 °C to water at 50.0 °? (Answer: 11274 J) 2) How much heat is required to raise the...

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5.2: The First Law of Thermodynamics.

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First law of thermodynamics: aka The Law of Conservation of Energy; a statement of our experience that energy is conserved in any process. We can express the first law in many ways. One of the more useful expressions is that the change in internal energy, ΔE , of a system in any process is equal to the heat, q , added to the system, plus the

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work, w, done ...

5.S: Thermochemistry (Summary) - Chemistry LibreTexts

LO MAPPED – The FRQ's that are listed on a worksheet marked as 'LO MAPPED', are mapped to SPECIFIC LO's in the May 2019 CED. It should be noted that NONE of the questions prior to 2020 were

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actually mapped by the College Board to the May 2019 CED, since they were all written prior to that document being produced.

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+ Cl₂ → 2 HCl + 1845 kJ (1) Is this reaction endothermic or exothermic?

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