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$4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$
4
a. What is the value of the coefficient x in this equation? 40.07 g/mol
b. What is

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the molar mass of C₃H₄? 2 mol O₂:1 mol H₂O c. What is the mole ratio of O₂ to H₂O in the above equation? 0.20 mol d. How many moles are in an 8.0 g sample of C₃H₄? 3z e. If z mol of C₃H₄ react, how many moles of ...

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Stoichiometry SECTION 3 PROBLEMS
Write the answer on the line to the left.
Show all your work in the space
provided. 1. 88% The actual yield of a
reaction is 22 g and the theoretical yield
is 25 g. Calculate the percentage yield.

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Section 1. Gases & Pressure. Pressure & Force. Pressure (P) is defined as the force per unit of area on a surface.
 $\text{pressure} = \text{force} \div \text{area}$ $P = F/A$.
Atmospheric pressure (atm) is the pressure exerted on an object due to the weight of the column of the air above it in the atmosphere. A .

Modern Chemistry Chapter 11 GASES

4? about 85°C b. What would be the

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boiling point of water if the air pressure over the liquid were reduced to 60 kPa? about 38 kPa c. What must the air pressure over CCl_4 be for it to boil at 50°C ? d. Although water has a lower molar mass than CCl_4 , it has a lower vapor pressure when measured? 20 STATES OF MATTER MODERN CHEMISTRY

10 States of Matter - Ms. Agostine's Chemistry Page

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